

**Chemistry Review
Science 10**

1. Give the compound name or formula for each of the following. Be careful, some may be multivalent or polyatomic!

NaCl	Sodium chloride
Lithium iodide	$\text{Li}^{+1} \text{I}^{-1} \rightarrow \text{LiI}$
Beryllium oxide	$\text{Be}^{2+} \text{O}^{2-} \rightarrow \text{BeO}$
Aluminum Sulfide	$\text{Al}^{+3} \text{S}^{-2} \rightarrow \text{Al}_2\text{S}_3$
Lead (II) nitrate ₊₂	$\text{Pb}^{+2} \text{NO}_3^{-} \rightarrow \text{Pb}(\text{NO}_3)_2$
$\text{Fe}_2(\text{SO}_4)_3$	Iron (III) sulfate
Sodium chlorate	$\text{Na}^{+1} \text{ClO}_3^{-1} \rightarrow \text{NaClO}_3$
Tin (II) sulfate ₊₂	$\text{Sn}^{2+} \text{SO}_4^{2-} \rightarrow \text{SnSO}_4$
Li_2CO_3	Lithium carbonate
Zinc sulfide	$\text{Zn}^{+2} \text{S}^{-2} \rightarrow \text{ZnS}$
Sodium fluoride	$\text{Na}^{+1} \text{F}^{-1} \rightarrow \text{NaF}$
Aluminum hydroxide	$\text{Al}^{+3} \text{OH}^{-1} \rightarrow \text{Al}(\text{OH})_3$
KMnO_4	Potassium permanganate
Ammonium sulfate	$\text{NH}_4^{+1} \text{SO}_4^{2-} \rightarrow (\text{NH}_4)_2\text{SO}_4$
Potassium oxide	$\text{K}^{+1} \text{O}^{-2} \rightarrow \text{K}_2\text{O}$

2. Give the compound name or formula for each of the following Molecular compounds.

Carbon tetrabromide	CBr_4
NI_3	Nitrogen triiodide
Oxygen difluoride	OF_2
SiCl_4	Silicon tetrachloride
Sulfur hexafluoride	SF_6
NO_2	Nitrogen dioxide
Diboron trisulfide	B_2S_3
PCl_5	Phosphorus pentachloride
Sulfur dioxide	SO_2